

LEACHING OF SECONDARY COPPER MINERALS USING REGENERATED FERRIC SULPHATE

C.J. FERRON — SGS

ABSTRACT

Secondary copper minerals such as chalcocite (Cu_2S) and covellite (CuS) are known to dissolve in acidic ferric sulphate solutions. However, the quantities of ferric sulphate required to achieve acceptable recoveries for high-grade copper feeds are prohibitive, unless the resulting ferrous sulphate is reoxidized and reused. A plant recently commissioned in Australia employs an autoclave for that purpose. This paper describes a novel approach, called the "Regenerated Ferric Sulphate Leaching (RFSL) process, in which the reoxidation of the ferrous ion is carried out under strictly atmospheric conditions, using SO_2/O_2 mixtures as the oxidizing medium. The process is very effective and rapid, and the reoxidation step can be accomplished during the leach itself or in a separate vessel. Several examples are presented for ores and concentrates originating from the USA, Central America and South East Asia. The overall concept of the process is also briefly assessed with respect to the main operating costs and sulfur balance.

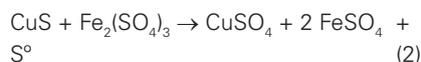
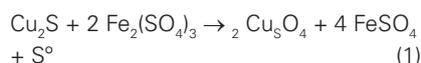
INTRODUCTION

Chalcopyrite, CuFeS_2 , is the most abundant copper sulphide mineral, and efficient treatment methods for producing high quality copper from chalcopyrite concentrates have been in use for several decades. The conventional treatment processes are pyrometallurgical and involve smelting, converting and electrorefining. Hydrometallurgical processes have also been developed and are waiting for commercial recognition. Very recently, a demonstration plant to produce 30,000 T Cu/year has been built by Phelps Dodge in Bagdad (Arizona), and will make use of high temperature pressure oxidation to process chalcopyrite concentrates. Secondary copper sulphide minerals such as chalcocite (Cu_2S) or covellite (CuS), although not as abundant as chalcopyrite, represent a non-negligible source of copper. The enrichment zone of porphyry copper deposits (e.g. Morenci), or the sulphidic zones of Central African copper deposits, are examples of secondary copper minerals deposits.

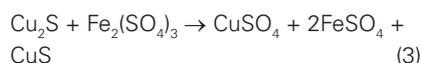
Secondary copper minerals concentrates are also typically smelted either directly or after sulphate roasting. Secondary copper sulphide minerals are normally less refractory than chalcopyrite and

as such appear more suitable to hydrometallurgical processes.

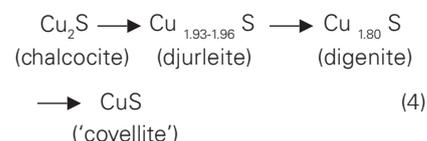
Acidified ferric sulphate is without a doubt the most common lixiviant for secondary copper minerals (1,2), and, even under autoclave conditions (3), it is believed that oxidation by ferric remains the most favoured mechanism. The dissolution reactions under atmospheric conditions can be written as follows for chalcocite and covellite, respectively:



In the 1930's already, it was recognized that the attack of chalcocite by ferric sulphate proceeded in two steps (4). The first step is fairly rapid and can overall be written as:



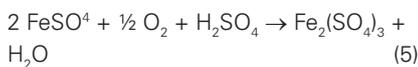
In fact, the reaction proceeds through the formation of a series of copper-deficient intermediates while cuprous ions diffuse through the solid to the surface, to be oxidized to cupric sulphate:



The second step, that can be written as for the naturally occurring covellite, is much slower, due to the fact that reactants and reaction products must diffuse through a sulphur layer. The presence of ligands, most particularly chloride, has been shown to significantly improve the kinetics of the chalcocite dissolution, increasing the rate of the first step (through the formation of soluble copper species not requiring oxidation of CuI to Cu^{II} , for example CuCl_2^-) and of the second step (through the formation of a porous crystalline sulfur layer) (5,6,7,8).

Two moles of ferric iron are required to dissolve one mole of copper, and therefore, for the process to be practical, the ferric iron has to be regenerated. There are different means available to the metallurgist to regenerate the ferric iron.

- Air or oxygen: From a thermodynamic point of view, oxygen is an oxidant strong enough to oxidize ferrous iron, but the kinetics of the oxidation reaction are normally slow in the acid range.
- Bacteria (for example thiobacillus ferro-oxidans): Biological oxidation of ferrous iron is a well known process and is commonly applied in heaps to treat low grade secondary copper ores (9). Tank bio-oxidation has also been proposed for secondary copper concentrates (10,11,12).
- Electro-oxidation. Anodic oxidation of ferrous was already proposed by the Inspiration Consolidated Copper Corporation in the 1930's (4).
- Pressure oxidation: Autoclaves have also been proposed as a means to oxidize ferrous iron, and under the higher temperatures and oxygen partial pressures involved, the reaction:



is kinetically fast.

Examples of the use of autoclaves to reoxidize ferrous iron to attack secondary copper minerals are the Dynatec Las Cruces process (13,14,15,16), the Mount Gordon Gunpowder process (17), and the Copper Range in situ process.

- SO_2/O_2 mixtures: The oxidizing properties of SO_2/O_2 mixtures to oxidize ferrous iron have been known since the 1920's work by the USBM. SO_2/O_2 mixtures have since been used to oxidize other ionic species such as CN- (Inco's process), or manganese, or arsenic. Improved processes for Ni/Co-Mn and Zn-Mn separation (18) resulted from these studies. Based on all these studies, regenerated ferric sulphate leach processes were proposed for copper sulphide minerals and for zinc sulphide minerals, whereby SO_2/O_2 mixtures were used to regenerate ferric sulphate, either in-situ (during the leach) or ex-situ (in a separate reactor). Initial

results have been presented for copper and zinc sulphide minerals (19,20).

EXPERIMENTAL

DESCRIPTION OF THE SAMPLES TESTED

Three samples were tested during this program; chemical and mineralogical analyses of these samples are presented in the following paragraphs. The first two samples were flotation concentrates, the first one produced during a pilot plant, the second one from a commercial operation in the South West USA. The last sample was an ore from a developing project.

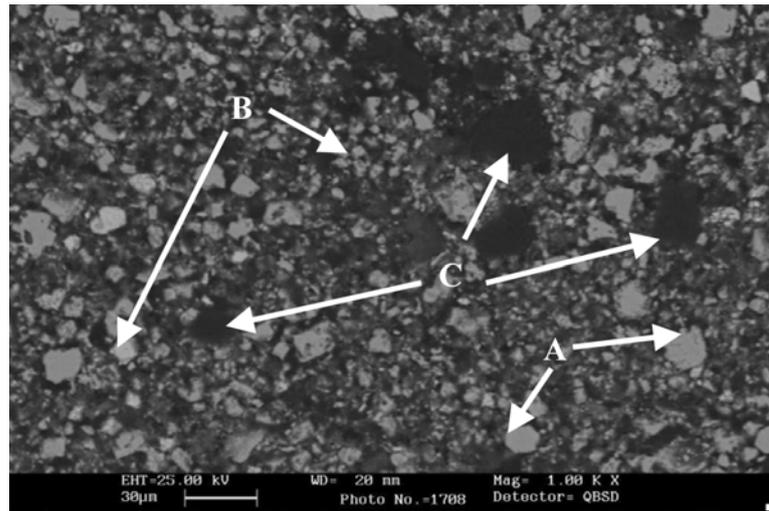


Figure 1 – Back-scattered electron (BSE) photomicrograph of Concentrate A illustrating the fine-grained nature of the sample, with abundant pyrite (A) and lesser amounts of covellite (B) and graphite (C).

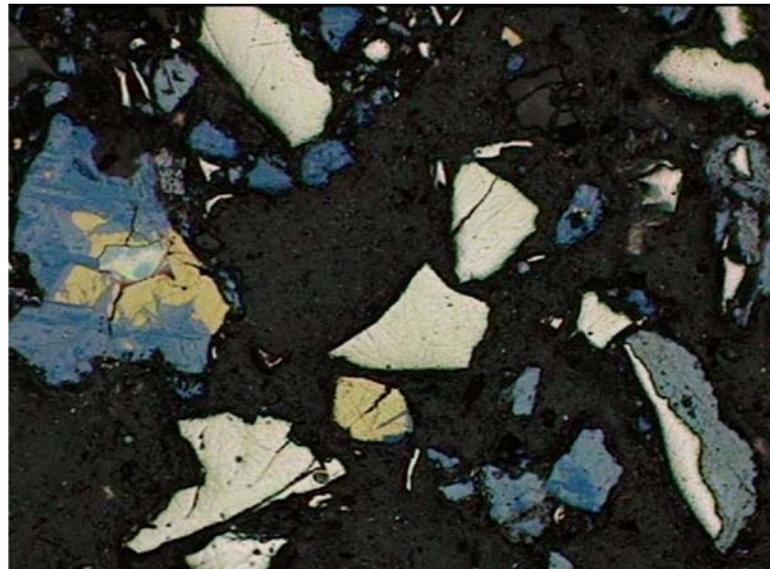


Figure 2 – Photomicrograph of Concentrate B: Reflected light; magnified 100x, field of view = 600 micrometers. Micrograph shows liberated pyrite, pyrite rimmed with chalcocite; liberated covellite-digenite grains, as well as covellitedigenite-chalcocopyrite middlings to the left.

Concentrate A was a flotation concentrate having pyrite as the main mineral and covellite as the major Cu-bearing mineral. A typical photomicrograph is presented in Figure 1. Concentrate B was also a flotation concentrate composed primarily of sulphides (94% weight), more particularly covellite (40% weight), chalcocite/digenite (18% weight) and pyrite (23% weight). A typical photomicrograph is presented in Figure 2. Concentrate C was an ore containing copper mostly as covellite, with minor chalcocopyrite and enargite. The gangue was composed primarily of quartz and pyrite.

RESULTS

CONCENTRATE A

Ferric Sulphate Leach Results

The initial tests examined the response of the concentrate to acidified ferric sulphate additions. Results are summarized in Table II.

The results are not surprising; at high temperatures (80-90°C), provided that sufficient ferric is present, between 91 and 96% of the copper is extracted in 3-4 hours. However, the resultant leach solution (~41 g/L Fe, ~23 g/L Cu) would pose certain challenges to be addressed.

The application of SO₂/O₂ mixtures during the leach (to regenerate in-situ the ferric needed to extract the copper) allows similar good extractions (>95% Cu) to be achieved with only 10 g/L iron in solution, as indicated in Table III. Such leach solutions would be much easier to process.

The Use of SO₂/Air Mixtures

The results presented above confirm the suitability of SO₂/O₂ mixtures to regenerate the ferric iron needed to maintain copper extractions. In some areas, the cost of oxygen could be prohibitive, and, therefore, there could be significant economic incentives in using SO₂/Air mixtures. Several tests were conducted to examine whether SO₂/Air mixtures could be used to regenerate the ferric iron. The results are presented graphically in Figure 3.

If air alone was used instead of SO₂/Air (1.5% SO₂ v/v) mixtures, the copper extraction at 70°C after 6 hours would be 80.5% as compared with 96.7% when using 1.5% v/v SO₂ in the air.

Table I: Chemical Analyses of the Samples Tested

CONCENTRATE	A	B	C
Origin	Caribbean	USA	Asia
%Cu	16.3	37.7	5.91
%Fe	24.1	17.1	11.2
%S ²⁻	36.7	29.2	8.6

(10% solids; concentrate as received (P₈₀ = 17µm))

TEST #	TEMP (°C)	DURATION (HRS)	INITIAL FE ³⁺ (g/L)	ACID ADDED (kg/t)	CU EXTRACTED (%)
1	80	3	10	305	56.3
2	80	3	40	305	90.9
3	90	4	40	305	96.1

Table II - Summary Results of Ferric Sulphate Leach

(10% solids, 2% SO₂ (v/v))

TEST #	TEMP (°C)	DURATION (Hrs)	INITIAL FE ³⁺ (g/L)	ACID ADDED (kg/t)	SO ₂ /O ₂	CU EXTRACTED (%)
1	70	6	10	305	no	57.3
4	60	6	10	305	yes	68.7
5	70	6	10	305	yes	96.1
6	90	6	10	305	yes	98.3

Table III - Summary Results of Regenerated Ferric Sulphate Leach

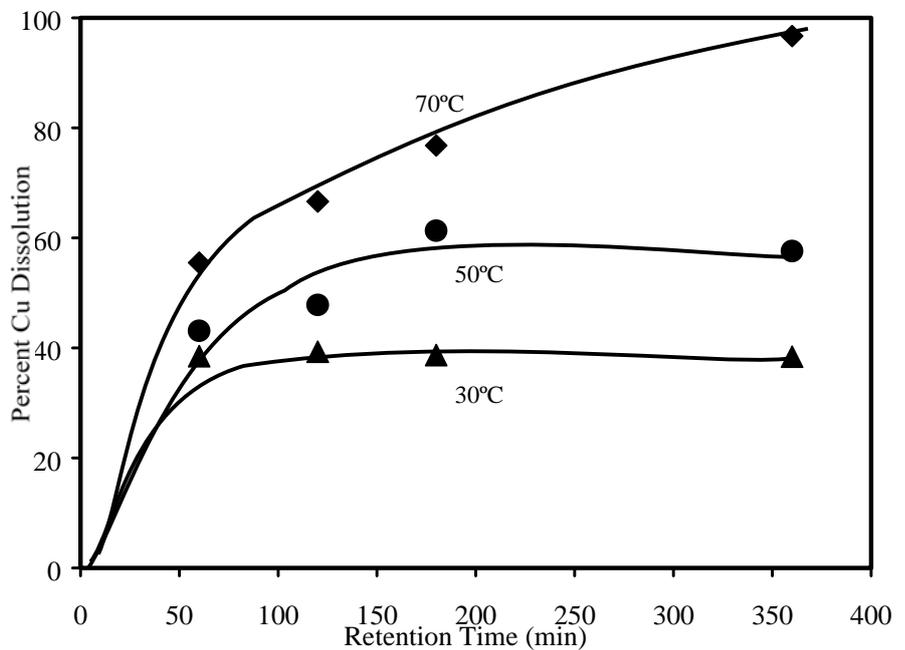


Figure 3 - Effect of Temperature on the Kinetics of Cu Leaching Using Regenerated Ferric Sulphate (SO₂/Air mixtures) (10% solids; feed as-received: P₈₀ = 17µm)

CONCENTRATE B

The initial series of tests examined the suitability of SO₂/O₂ mixtures to regenerate ferric during the leach. The results are presented in Figure 4. Without regeneration, the best results that can be expected are 37.7% Cu dissolved, when all the ferric has been consumed. Adding oxygen during the leach promotes ferric regeneration, and, therefore, copper dissolution continues and achieves 62% in 6 hours under the conditions of the test. Adding SO₂ (2%) and replacing oxygen with air, while keeping the same flow of gases, produces the same copper extraction in 6 hours. Adding SO₂ to oxygen (5% SO₂ v/v), while keeping the same flow of gases, significantly improves the copper dissolution that reaches 88% in 6 hours. Clearly, SO₂/O₂ (or even SO₂/air) mixtures are more efficient than oxygen alone to regenerate ferric. The procedure is also quite reproducible (at the laboratory scale), as indicated in Figure 5. The effect of varying the proportion of SO₂ in the SO₂/O₂ mixture is presented in Figure 6.

The results indicate that 2% (v/v) SO₂ was sufficient to induce the maximum benefits of the gas blends. Higher percentages did not improve the results further, and there is some indication that 10% SO₂ in the blend had started reversing the benefits (too reducing conditions), at least initially.

As could be extrapolated from the shape of the dissolution curves, the dissolution was not completed in 6 hours, and longer retention times would increase the copper yield. This was confirmed as illustrated in Figure 7.

As expected for the dissolution of copper from covellite when using acidified ferric sulphate, the rate-limiting step is the migration of reactants and/or reaction products through the sulfur layer. This was confirmed by the linear correlation (at least at the beginning) of the type: g/L Cu dissolved = $\kappa \sqrt{t}$, as illustrated in Figure 8.

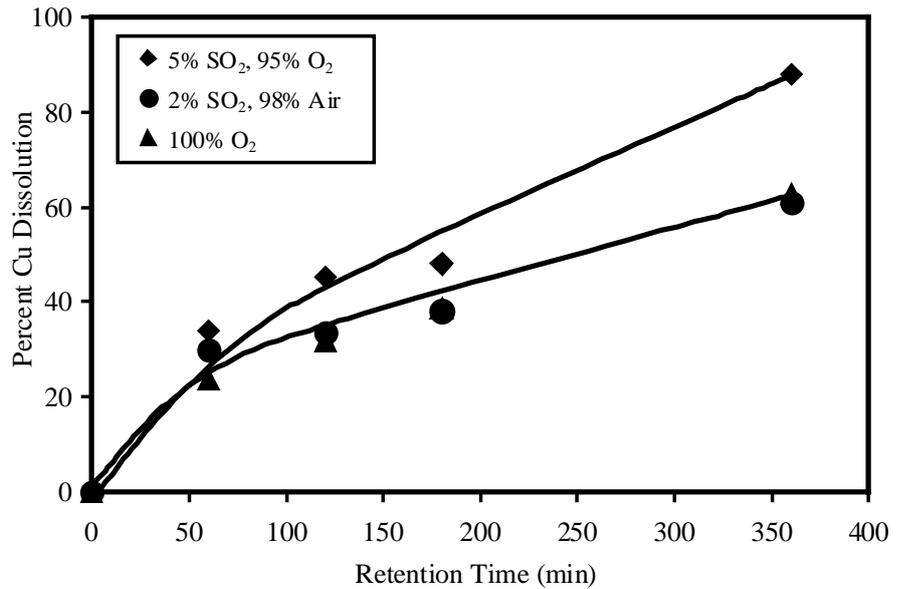


Figure 4 - Effect of SO₂ Addition on the Regeneration Efficiency (80°C, 4% solids, P₈₀ = 55 µm, 10g/L Fe³⁺)

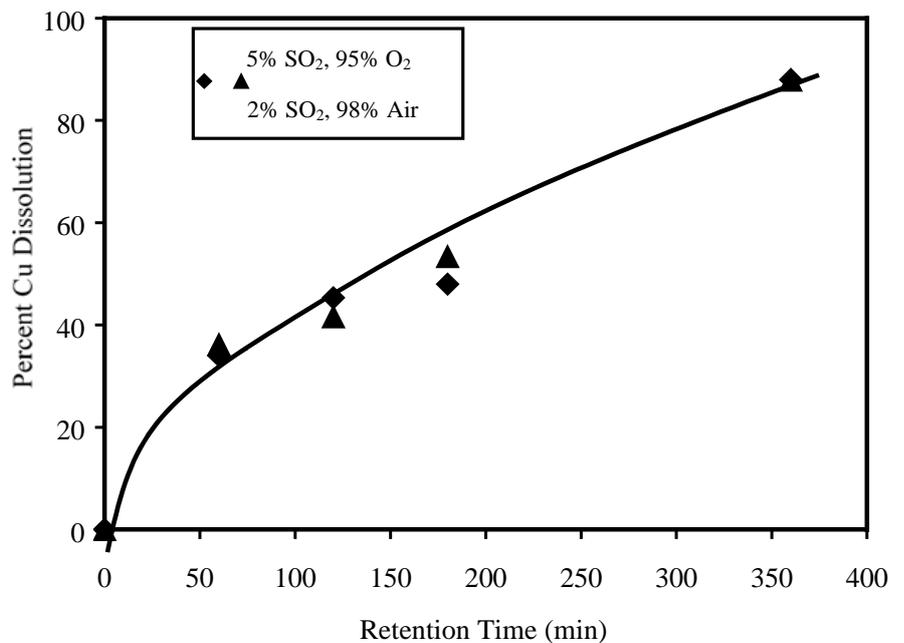


Figure 5 - Reproducibility of Experimental Procedure

CONCENTRATE C

Initial tests were carried out to confirm the beneficial effect of SO₂/O₂ mixtures to improve the regeneration of ferric during the leach (Figure 9). The results confirmed the beneficial effect of SO₂/O₂ mixtures for the regeneration of ferric and the improvement on copper dissolution. Without in-situ regeneration,

30 g/L Fe³⁺ must be added to achieve acceptable copper extraction (~95% in 4 hours). With ferric regeneration (using SO₂/O₂ mixtures with 2% SO₂), the required ferric concentrate concentration could be reduced to 10, 5 and even 2 g/L to achieve similar extractions. Oxygen alone is not as efficient to regenerate ferric as illustrated in Figure 10.

Comparing the results of various tests, one can easily see that the addition of small amounts of SO₂ to the oxygen (same flowrate) resulted in a largely improved ferric regeneration, with 88% of the copper dissolved in 60 minutes compared to 80% in 120 minutes when pure oxygen was used. With slightly higher ferric (5 g/L), SO₂/O₂ mixtures led to 93% Cu dissolution in only 60 minutes.

DISCUSSION

The results presented here on three samples (containing chalcocite or covellite minerals primarily as copper-bearing minerals) certainly confirmed the results presented earlier, namely that SO₂/O₂ mixtures are a very effective oxidant to regenerate ferric sulphate solutions. All the results presented here deal with the in-situ regeneration of ferric sulphate (i.e. the regeneration of ferric sulphate is accomplished in the leach reaction vessel while copper sulphide minerals are being attacked), but earlier results indicated that the regeneration could easily be accomplished on solutions ex-situ in a separate vessel.

Examination of the mechanism(s) of ferric sulphate dissolution of chalcocite and covellite was not the primary objective of this work, since that issue had been tackled by numerous researchers.

To understand the reason for such an improvement in the efficiency of copper dissolution when regenerating ferric sulphate with SO₂/O₂ mixtures, one has to examine what part of the reaction path necessitates high ferric concentrations.

According to a 1974 review (1), the rate-limiting step for chalcocite dissolution when using acidified ferric sulphate is the diffusion in solid solution of ferric ions to the Cu₂S surface; at the surface of the chalcocite, the oxidation by ferric of Cu(I) migrated from the inner mineral must therefore occur rapidly; when all the mineral chalcocite has been transformed to blue-remaining covellite ('blaubleibender'), ferric has to diffuse through the sulfur layer

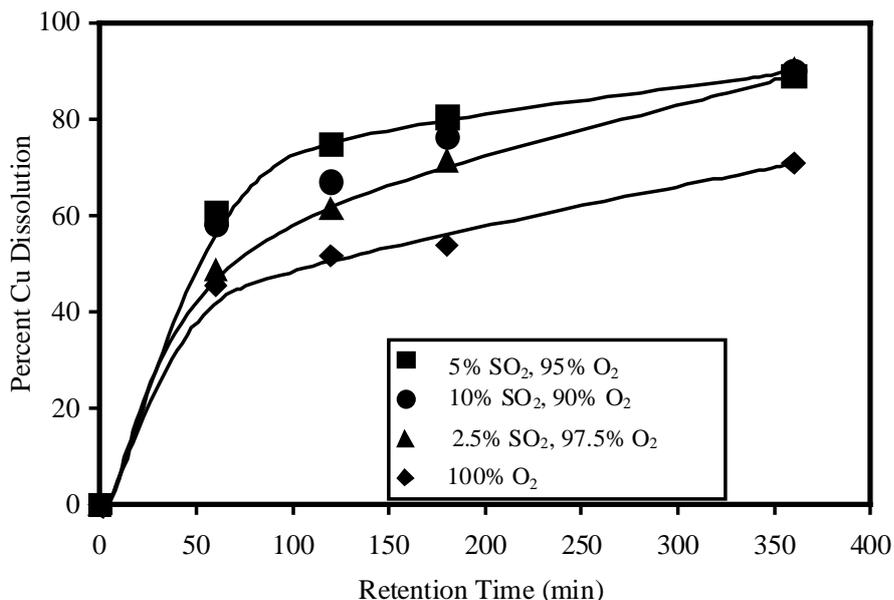


Figure 6 - Effect of SO₂ to O₂ Ratio on Cu Dissolution (70°C, 2% solids, P₈₀ = 13 µm)

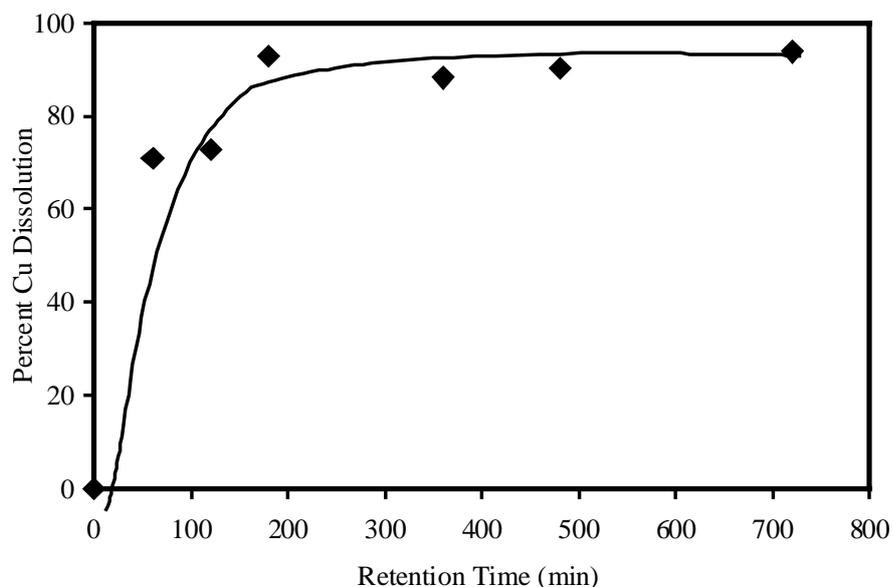


Figure 7 - Ultimate Extraction (70°C, 2% solids, P₈₀ = 13 µm)

to the inner mineral, while ferrous sulphate and copper sulphate have to diffuse through the same sulfur layer to the outer mineral. The relatively low activation energies measured by several authors (1.5 to 7 Kcal/mole, 6 to 18 kJ/mole) indicate that the sulfur layer produced from 'blaubleibender' must be quite porous and does not constitute a high barrier to the progress of the reaction. It can therefore be postulated that the contribution of SO₂/O₂ mixtures to the improved kinetics of chalcocite dissolution is to maintain a higher concentration of ferric in solution,

thereby favouring Fe³⁺ diffusion because of the concentration gradient.

The same reviewer for covellite dissolution using acidified ferric sulphate solution concluded to a chemically controlled process, with a significantly higher activation energy (8-25 kcal/mole, 32-100 kJ/mole). Some results presented here point to a mechanism controlled by diffusion through a reaction product when using a SO₂/air mixture. Additional work would be required to fully understand the reaction mechanism.

A rate expression for the oxidation of ferrous solution when using SO₂/O₂ mixtures has been presented in earlier publications and can be summarised as:

$$\frac{d\text{Fe}^{2+}}{dt} = k' \frac{[\text{SO}_2]^a [\text{SO}_3]^{1-b}}{[\text{H}_2\text{SO}_4]^c} \exp \frac{-E}{RT} \quad (6)$$

where a, b and c are 2.1, 1.1 and 2.5 respectively, and E equals 6,650 J/mole and k' = k''[pO₂].

It is unquestionable that the SO₂/O₂ mixtures enhance the kinetics of ferric sulphate leaching of chalcocite and covellite ores or concentrates. To assess the industrial applicability of a new process, one cannot stop to that consideration and must also consider the implications of introducing SO₂ on the overall process sulfur balance.

The first point to consider is that little pyrite is attacked by ferric sulphate, and therefore, sulphate ions introduced by the oxidation of pyrite are a minor contributor to the overall solution sulphate balance. Moreover, there is no evidence that pyrite oxidation by ferric sulphate is increased when adding SO₂/O₂ mixtures.

If one examines the overall sulphate balance of the whole process, including leaching, ferric regeneration, solvent extraction and copper electrowinning when using oxygen alone or SO₂/O₂ mixtures, one arrives at the following reactions:

a) For Chalcocite

With oxygen :

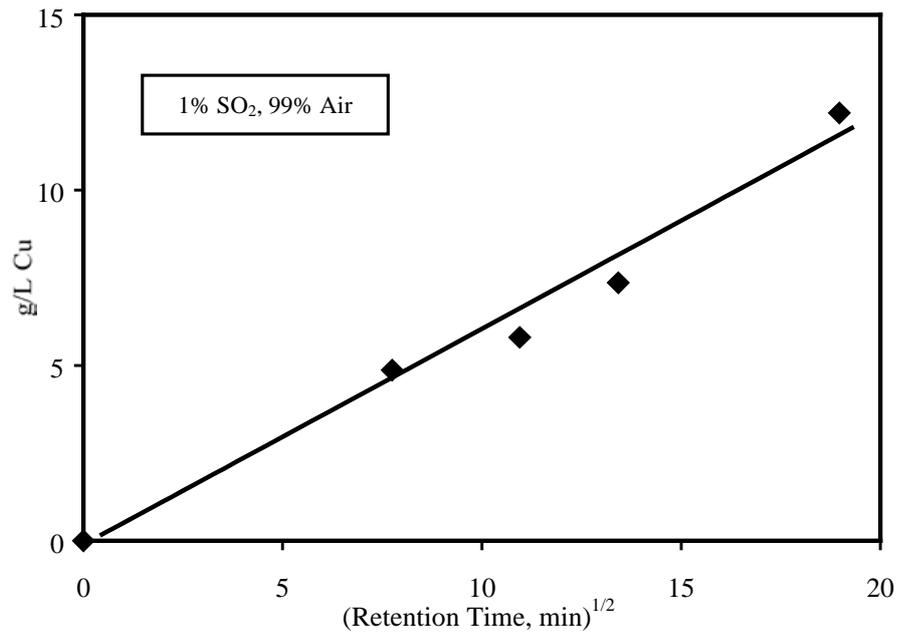
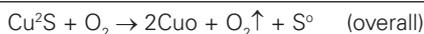
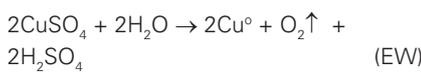
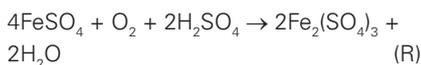
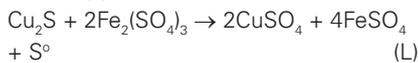


Figure 8 - Kinetics of Cu Dissolution from Concentrate B (80°C, 4% solids, P₈₀ = 55 µm)

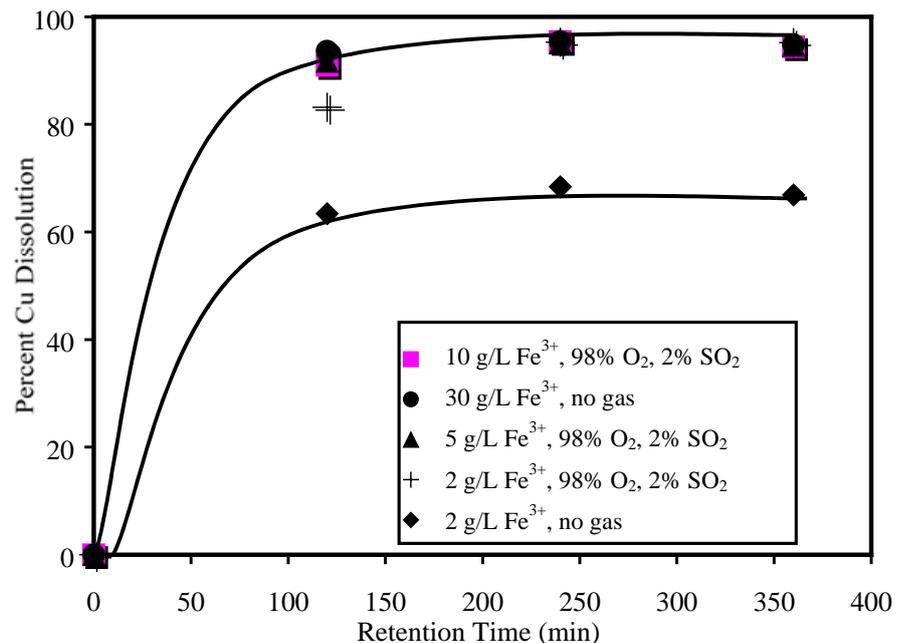
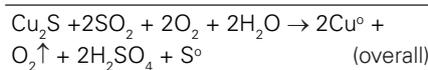
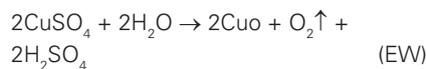
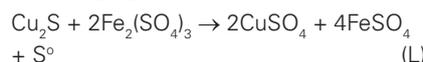


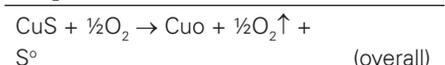
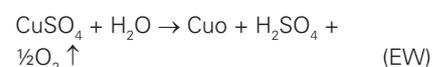
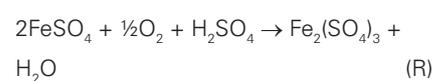
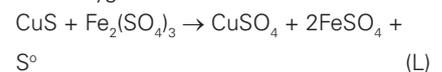
Figure 9 - Effect of SO₂/O₂ Mixtures on Cu Extraction (80°C, 10% solids, no regrind)

With SO₂/O₂ :

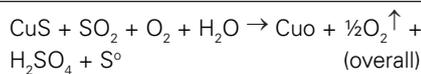
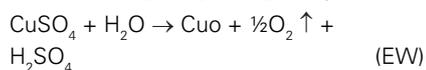
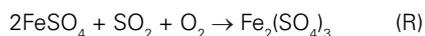
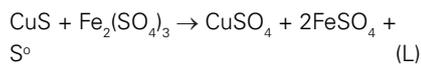


b) For Covellite

With oxygen :



With SO_2/O_2 :



where L, R, EW represent the leaching, the regeneration and the electrowinning stages, respectively.

Based on the overall reactions, when using SO_2/O_2 mixtures, the theoretical consumptions are 1 mole SO_2 /mole of Cu and 1 mole O_2 /mole of Cu. Assuming a cost of 30\$/T of S and 30\$/T of O_2 , these consumptions amount to 0.34 US cent/lb of Cu and 0.69 US cents/lb of Cu, respectively.

The coefficient of utilisation for SO_2 in this type of process is excellent, and 10% excess over stoichiometry is quite sufficient. Oxygen is used in large excess and the unused portion should be recycled. Alternatively, air could be used instead of pure oxygen; results presented here indicate that SO_2 /air mixtures are at least as efficient as oxygen alone, but the kinetics would be slower with oxygen alone.

The overall reactions also indicate that one mole of H_2SO_4 is produced per mole of copper dissolved, corresponding to 0.70 kg H_2SO_4 /lb Cu. Some of this acid will be consumed by the acid-consuming gangue present with the concentrate, and this will obviously vary from concentrate to concentrate. The excess acid could be used to attack copper oxide concentrates if available (for example in Central Africa) or used on a heap. The worst case situation would be that all that excess acid would have to be neutralised using limestone, and this would represent a cost of 2.1 US cent/lb Cu (assuming a limestone cost of 30 US\$/T limestone).

CONCLUSIONS

The regenerated ferric sulphate leach process (RFSL) as described in this paper and in a previous presentation is undoubtedly an efficient means to treat secondary copper (covellite + chalcocite) concentrates or ores: under mild conditions (70°C, no overpressure, pH ~1.5-2.0), in a few hours, acceptable copper extractions ($\geq 95\%$) can be achieved without having to resort to autoclaves. The produced PLS has a suitable chemical composition (acid, Cu, Fe) so that it can directly be fed to conventional solvent extraction circuits.

The equipment to be used consists of conventional and simple tanks, mixers, pumps, thickeners/filters, gas spargers; since no chloride is used in the process, corrosion is not a significant issue. Because of the process simplicity, it can be safely scaled up from small-scale operations to larger commercial plants.

Regarding operating costs, it is not possible to give a figure because each case will be different, depending on the concentrate to be treated, its acid-consuming components, the availability of SO_2 (smelter, elemental sulfur), the cost of oxygen (electric power), and the availability of cheap limestone or a source of oxide copper to consume excess acid (oxide copper concentrate, heap leach operations).

To be able to match the composition of the resulting PLS to the requirement of conventional solvent extraction circuits, the copper tenor has to be adjusted to less than about 20 g/L, and therefore, the percent solids in the leach circuit has to be adjusted accordingly. In practice, this indicates that the process is better suited to high grade ores or low grade concentrates (for example 10-15% Cu). This brings an additional benefit from this process since it allows the mineral processor to increase the overall copper yield by producing a lower grade concentrate, for example a rougher or a first cleaner concentrate.

No detailed heat balance has been prepared during this bench scale test program since this is much more precisely carried out during a pilot plant program, but indications from the literature are that the process would not require the addition of external heat.

Finally, concerning the recovery of precious metals, if present in the original concentrate, they end up in the RFSL residue together with elemental sulfur, the basically untouched pyrite and all the insolubles. Their recovery is therefore a challenge, but there are several alternatives that have been considered, although at an early stage.

REFERENCES

1. J.E. Dutrizac, R.J. MacDonald, "Ferric Ion as a Leaching Medium" *Minerals Sci Eng*, Vol 6, No. 2, April 1974, pp59-100.
2. R.J. Roman, B.R. Benner, "The Dissolution of Copper Concentrates" *Minerals Sci. Eng.*, Vol 5, No. 1, Jan 1973, p3-24.
3. M.H. Mao, E. Peeters, "Acid Pressure Leaching of Chalcocite" in *Hydrometallurgy – Research, Development and Plant Practice*, Ed. K. Osseo-Asare, J. Miller, TMS 1983, Atlanta, Georgia, USA, pp243-260.
4. J.D. Sullivan, "Chemistry of Leaching Chalcocite" *USBM Technical Paper 473*, 1930.
5. W.W. Fisher, F.A. Flores, J.A. Henderson, "Comparison of Aqueous Chalcocite Dissolution in the Oxygenated, Aqueous Sulphate and Chloride Systems", *Minerals Engineering*, Vol 5, No. 7, 1992, pp817-834.
6. W.W. Fisher, "Comparison of Chalcocite Dissolution in the Sulphate, Perchlorate, Nitrate Chloride, Ammonia and Cyanide Systems", *Minerals Engineering*, Vol 7, No. 1, 1994, pp99-103.
7. C.Y. Cheng, F. Lawson, "The Kinetics of Leaching Chalcocite in Acidic Oxygenated Sulphate-Chloride Solutions" *Hydrometallurgy* 27, 1991, pp249-268.

8. C.Y. Cheng, F. Lawson, "The Kinetics of Leaching Covellite in Acidic Oxygenated Sulphate-Chloride Solutions" *Hydrometallurgy* 27, 1991, pp269-284.
9. S. Bustos, R. Espejo, C. Gonzalez, R.E. Scheffel, "Bacterial Heap Leaching of Covellite" *Randol Copper 2000*, Ed. H. Von Michaelis, Randol International, Vancouver, Canada.
10. A. Bruynesteyn, "Microbiological Leaching-Research to Date and Future Application" AIME meeting, Denver, Colorado, USA, February 1970.
11. A. Bruynesteyn, D.W. Duncan, "Microbiological Leaching of Sulphide Concentrates" *Canadian Metallurgical Quarterly*, Vol 10, No. 1, 1970, pp57-63.
12. J.M. Weston, "Continuous Biological Leaching of Copper from a Chalcocite Ore and Concentrate in a Saline Environment" Master Thesis, University of British Columbia, July 1995.
13. M.J. Collins, K. Buban, F. de Kock, R. Kalanchey, T. Xue, "Copper Processing in the Dynatec Miniplant" *Alta Copper 1998*, Ed. A. Taylor, Alta Metallurgical Services, Brisbane, Australia.
14. K. Norris, M. Makwana, "The Las Cruces Copper Plant Process Engineering Design" XI International Congress of Industry, Mining and Metallurgy, Zaragoza, Spain, June 2002.
15. R.M. Berezowsky, T. Xue, M.J. Collins, I. Barton-Jones, M. Southgate, J.K. MacLean, "Pressure Leaching Las Cruces Copper Ore" *JOM*, Dec 1999, pp36-40.
16. N. Smalley, G. Davis, "Operation of the Las Cruces Ferric Sulphate Leach Pilot Plant" *Minerals Engineering*, Vol 13, No. 6, 2000, pp559-608.
17. G. Richmond, "The Mt. Gordon Copper Autoclave Process of Western Metals Limited" *Alta Copper 2001*, Ed. A. Taylor, Alta Metallurgical Services, Brisbane, Australia.
18. C.J. Ferron, "The Control of Manganese in Acidic Leach Liquors, with Special Emphasis to Laterite Leach Liquors" *Nickel/Cobalt Alta Conference*, Ed. A. Taylor, Alta Metallurgical Services, Perth, Australia, May 2002.
19. C.J. Ferron, "New Atmospheric Leach Process for Copper Sulphide Ores and Concentrates" *Copper 99*, Ed. S.K. Young, D.B. Dreisinger, R.P. Hackl, D.G. Dixon, Vol. IV, TMS, Phoenix, Arizona, USA, Oct. 1999
20. C.J. Ferron, "Atmospheric Leaching of Zinc Sulphide Concentrations using Regenerated Ferric Sulphate Solutions" in *Lead-Zinc 2000*, Ed. J.E. Dutrizac, J.A. Gonzales, D.M. Henke, S.E. James, A.H.J. Siegmund, TMS Fall Extraction & Process Metallurgy Meeting, Pittsburgh, USA, Oct 2000.

CONTACT INFORMATION

Email us at minerals@sgs.com

WWW.SGS.COM/MINERALS

© 2011 SGS. All rights reserved. The information contained herein is provided "as is" and SGS does not warrant that it will be error-free or will meet any particular criteria of performance or quality. Do not quote or refer any information herein without SGS' prior written consent. Any unauthorized alteration, forgery or falsification of the content or appearance of this document is unlawful and offenders may be prosecuted to the fullest extent of the law.